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(54) **Treating subterranean formations with cellulose and guar derivatives**

Behandlung von unterirdischen Lagerstätten mit Zellulose- und Guaderivaten

Traitement de formations souterraines à l'aide de dérivés de cellulose et de guar

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D r i p t i o n

[0001] This invention relates to a method of treating a subterranean formation with certain derivatives of guar and cellulose.

[0002] Petroleum recovery operations, such as well stimulation and gravel packing, often require the use of fluid compositions capable of suspending particles. In gravel packing operations, a pack of gravel is placed on the exterior of a perforated or slotted liner or screen which is positioned across an unconsolidated formation. The resulting structure presents a barrier to migrating sand from the formation while still permitting fluid flow. The gravel is carried to the formation in the form of a slurry by mixing gravel with a viscosified fluid. Once the gravel is placed in the wellbore, the viscosified carrier fluid is degraded and returned to the surface.

[0003] Treating fluids are used similarly in stimulating subterranean formations. The viscosified fluid carries a propping agent through the wellbore and into both natural fractures and fractures in the formation caused by hydraulic pressure. Once the desired fracture occurs, the fluid is degraded and returned to the surface leaving the proppant in the formation to provide a conductive channel through which formation fluids can flow. In both stimulation and gravel packing operations the most desirable treating fluid reduces friction pressure as the fluid is pumped through the tubular goods and transports the propping agent or gravel to the formation without particle settling in the wellbore during placement.

[0004] Treating fluids with these properties are generally comprised of a hydratable polysaccharide, including but not limited to guar, guar derivatives and cellulose derivatives. These polymers viscosify aqueous liquids to form solutions which inhibit particle settling to a limited extent by virtue of viscosity. However, these polymer solutions can approach near zero particle settling rates upon crosslinking with multivalent metal cations to form highly viscoelastic gels. The utility of these gels is well known in the art of petroleum recovery operations.

[0005] Cellulose derivatives are the preferred viscosifying polymers for certain petroleum recovery operations because they degrade to lose viscosity without generating water insoluble particles or residue. The water insoluble particles are believed to remain in the formation and may cause formation plugging or impair the permeability of sand or gravel packs. However, cellulose derivatives have had limited use in many petroleum applications because most derivatives are either salt sensitive or not crosslinkable. Non-ionic derivatives of cellulose are generally not crosslinkable because the polymer lacks a site for attachment of a multivalent metal cation. Examples of this type include hydroxyalkyl cellulose ethers, methyl cellulose, ethyl cellulose, and hydroxyalkyl methyl cellulose. A crosslinkable non-ionic cellulose derivative has been prepared and described in U.S. patent specification nos. 4,523,010 and 4,552,215 to which reference should be made. In these specifications, dihydroxypropyl hydroxyalkyl cellulose is prepared by a condensation reaction of glycidol with hydroxyethyl cellulose under alkaline conditions. The glycidol addition along the HEC polymer chain provides a site of attachment for multivalent metal cations.

[0006] Anionic cellulose derivatives are normally substituted with carboxyl groups along the polymer chain. The carboxyl groups complex with polyvalent metal cations, such as aluminum. Gels formed with this chemistry tend to have limited structural stability at formation temperatures of about 250°F (121°C). In addition, carboxylate substituents render the polymer salt sensitive, i.e. the viscosity of the polymer in a salt solution is less than the viscosity in water. Salt sensitivity is not a desirable property because the aqueous liquids used in recovery operations most generally contain chloride salts to inhibit the swelling of formation clays.

[0007] We have now devised a new method of effecting crosslinking in selected graft copolymers of cellulose derivatives, which are generally non-ionic in character. Methods of grafting monomers on polyhydroxy containing compounds are well known in the art. The process is described in U.S. patent specification no. 2,922,768 to which reference should be made, U.S. patent specification nos. 4,982,783, 5,067,565 and 5,122,549, to which reference should also be made, describe processes by which crosslinkable cellulose derivatives are prepared by grafting vinyl or allyl monomers having a crosslinkable substituent onto the cellulose derivative. The resulting copolymer is non-ionic and crosslinks readily with polyvalent metal cations to form stable viscoelastic gels.

[0008] The surprising discovery now has been made that certain graft copolymers of hydroxyethyl or hydroxypropyl cellulose, prepared by a redox reaction with vinyl phosphonic acid monomers or polymers and hydroxyethyl or hydroxypropyl cellulose, can be crosslinked by the addition of a Lewis base or Bronsted-Lowry base or mixture of such bases to an aqueous solution, which contains at least a trace amount of at least one divalent cation, containing the graft copolymer. The base utilized generally is substantially free of polyvalent metal ions that is, metal ions having more than one valence state. The crosslinked gel formed by the addition of the base can be broken by the addition of an acid generating compound or other conventional breakers. The crosslinked gels are particularly useful in petroleum recovery operations in that the crosslinked cellulose derivatives provide highly viscous gels that can be degraded without generating significant quantities of water insoluble residue. The discovery also has been made that guar and hydroxypropyl guar can be grafted by the described technique and caused to form crosslinked gels upon the addition of a Lewis base or Bronsted-Lowry base.

[0009] The invention provides a method of treating a subterranean formation, which method comprises contacting

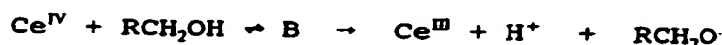
the formation with a treating fluid composition which comprises a mixture of (i) an aqueous liquid containing at least a trace amount of at least one divalent cation, (ii) a polymer derivative made by reacting one or more of the polymers guar, hydroxypropylguar, hydroxyethyl cellulose and hydroxypropyl cellulose, with a vinyl phosphonic acid in the presence of a redox system, and (iii) a crosslinking activator comprising a Lewis base or a Bronsted-Lowry base which is substantially free of polyvalent metal ions. The composition may contain a particulate agent if desired.

[0010] The present invention also provides a method of crosslinking an aqueous solution of a polymer derivative comprising a graft copolymer of hydroxyethyl cellulose, hydroxypropyl cellulose, guar or hydroxypropyl guar.

[0011] The fluid composition can be used in petroleum recovery operations, such as stimulation, gravel packing and other well completion operations such as temporary plugging agents and the like. In these operations, the treating fluid performs a variety of functions, for example, a highly viscoelastic fluid is often times required to transport propping agents or gravel packing materials to the formation without settling. In addition, the treating fluid must have stable viscosity at formation temperatures. The present invention provides such a fluid.

[0012] An aqueous liquid is used to solubilize the novel copolymer of the present invention. The term "aqueous liquid" is used hereafter to mean any liquid containing sufficient water to at least partially hydrate the copolymer and result in an increase in the viscosity of the fluid. Aqueous liquids used in petroleum recovery operations normally contain sodium chloride, potassium chloride, calcium chloride, sodium bromide and other bromides, tetramethylammonium chloride or the like to weight the fluid or inhibit the swelling of clays generally found in subterranean formations. The pH of the aqueous liquid must be compatible with the selected crosslinking agent and must not adversely affect the hydration of the copolymer. Such liquids generally contain at least a trace amount of at least one divalent cation, generally in the form of contaminants in the aqueous liquid.

[0013] In one embodiment of the present invention, the crosslinkable copolymers are prepared by reacting certain vinyl monomers comprising vinyl phosphonic acid (VPA), with a cellulose derivative using a redox system comprising ceric ions and nitric acid. The generalized reaction is believed to be represented by the formula:



where B is the ceric-alcohol complex, RCH_2OH is the cellulose derivative and $\text{RCH}_2\text{O}^\cdot$ is a free radical. Graft copolymerizations of cellulose commonly use chemical initiators, such as ceric ions. In acid media, ceric ions oxidize 1,2-glycols with the formation of a free radical on a reducing agent, which is the cellulose derivative in this case. The free radical produced on the cellulose derivative initiates polymerization with the vinyl group of the monomer to produce the graft copolymer.

[0014] The cellulose derivative of this invention is preferably a hydroxyalkyl cellulose having a hydroxyalkyl molar substitution from about 1.5 to about 3.0. Molar substitution is defined as the average number of moles of a substituent group present per anhydroglucose unit of the cellulose material. The alkyl group is selected from the group of ethyl, propyl and mixtures thereof. The preferred hydroxyalkyl cellulose is hydroxyethyl cellulose (HEC) having a molar substitution in the range of about 1.8 to about 2.5. Preferably in this invention, the hydroxyalkylation of the cellulose is preformed in a separate reaction. Hydroxyethyl cellulose is usually formed by reacting ethylene oxide with cellulose under extreme alkaline conditions and is available commercially.

[0015] The copolymers of the present invention are rendered crosslinkable by grafting monomers comprising a vinyl phosphonic acid to the cellulose derivative. The monomers have the reactive $\text{CH}_2=\text{C}-$ moiety that is believed to enable the monomer to attach to the cellulose derivative.

[0016] Typically, graft copolymerizations are carried out in aqueous media wherein the polymer is dissolved or dispersed. Copolymers of this invention were prepared in acetone (55% to 90%) and water (45% to 10%) or methanol (about 70%) and water (about 30%). Reactions were carried out in a 1 liter kettle with a stirrer or a 1 liter jar at about 20°C to about 60°C. The ratio of cellulose derivative to aqueous medium ranges from about 1 gram per 100 ml to about 1 gram per 2 ml. The preferred ratio is from about 1 gram per 6 ml to 1 gram per 4 ml. The ratio of cellulose derivative to grafting VPA monomer ranges from about 3 gram per 1 ml to about 25 gram per 1 ml. The preferred ratio is from about 6 gram per 1 ml to about 16 gram per 1 ml.

[0017] The polymerization reaction of the present invention may be chemically initiated by a redox system comprising ceric ions in acidic medium. Ceric ions may be provided, for example, by salts such as ceric nitrate, ceric sulfate, ceric ammonium nitrate, and ceric ammonium sulfate. The preferred ceric initiator of the present invention is a solution of ceric ammonium nitrate in 1N nitric acid. Ceric ammonium nitrate is present in an amount of from about 0.00075 mole per 100 ml to about 0.005 mole per 100 ml reaction medium.

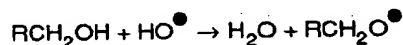
[0018] The ceric initiator may be added slowly to the reaction material over a time period of about 30 to 90 seconds or longer. Reaction times vary from about 10 minutes to 20 hours depending on reaction conditions or the particular grafting monomer. Grafting reaction efficiency is generally less than about 50%. After the reaction is complete, the

polymerization product is washed with acetone, filtered, and dried.

[0019] In another embodiment of the present invention, the crosslinkable copolymers are prepared by reacting certain vinyl monomers having a crosslinkable substituent with a cellulose derivative using a redox system comprising the reaction product of hydrogen peroxide with a ferrous salt. The generalized redox reaction is believed to be represented by the formula:



and the generalized initiation reaction is believed to be represented by the general formula:



[0020] An advantage of this initiator is that radical production occurs at a reasonable rate over a wide temperature range whereby reactions can be carried out at room temperature, if desired. The free radical produced on the cellulose derivative initiates polymerization with the vinyl group of the monomer to produce the graft copolymer.

[0021] Typically, the graft copolymerization is carried out in aqueous media wherein the polymer is partially dissolved or dispersed. Copolymers were prepared in acetone/water mixtures containing from about 55 to about 90% acetone.

Reactions were carried out in a 1 liter kettle with a stirrer or a 1 liter jar at a temperature of from about 20 to about 60°C. The ratio of cellulose derivative to aqueous medium ranges from about 1 gram per 100 ml. to about 1 gram per 2 ml. The preferred ratio is from about 1 gram per 2 to 5 ml. The ratio of cellulose derivative to grafting VPA monomer ranges from about 5 to about 40 grams per 1 gram of monomer. The preferred ratio is from about 6 to about 16. It is to be understood that the ranges set forth above are merely exemplary and that other temperatures, concentrations and the like may be utilized to prepare the reaction product.

[0022] The polymerization reaction of this embodiment of the invention is chemically initiated by a redox system comprising the reaction product of hydrogen peroxide with a ferrous salt. Ferrous ions may be provided, for example, by salts such as ferrous ammonium sulfate, ferrous chloride, ferrous sulfate, ferrous acetate, ferrous oxalate, ferrous acetylacetonate and the like. A preferred source of ferrous ions is ferrous ammonium sulfate. Alternatively, other commonly used metal ion reductants may be utilized in place of the ferrous ions to generate the free radicals necessary to effect grafting and other forms of hydrogen peroxide such as t-butylhydroperoxide may be used.

[0023] The initiator may be added slowly to the reaction material over a time period of about 30 to 90 seconds or longer. Reaction times vary from about 15 minutes to about 4 hours depending upon the reaction conditions or the particular grafting monomer. Grafting reaction efficiency (% of monomer grafted) is generally less than about 75%.

After the reaction is complete, the polymerization product is washed with acetone, filtered and dried.

[0024] In a preferred method of effecting the graft copolymerization, the grafted polymer product is retained in a substantially storage stable slurry form. Typically, the media comprises a polyglycol, such as polypropylene glycol having molecular weights up to about 1000 such as PPG-250 to PPG-1000 from Texaco Chemical Co., various polyethylene glycols and homopolymers of 1,2 butylene oxide having a molecular weight of from about 200 to about 400 which are present in an amount of from about 70 to about 95 percent by weight of the media and the remainder generally being water. The media also may comprise tetramethylammonium chloride in a similar amount or in admixture with a polyglycol. In a preferred embodiment the polyglycol comprises from about 86 to 92 percent by weight of the media. Reactions were carried out in a 5 liter kettle with a stirrer at a temperature of from about 20 to 60°C. The ratio of cellulose derivative to media ranges from about 1 gram per 100 ml to about 1 gram per 2 ml. The preferred ratio is from about 1 gram per 2 to 5 ml. The reaction media also may include a quantity of a dispersant or thixotrope such as alkyl quaternary ammonium montmorillonite (Claytone AF from E.C.C. America, Inc.) or dimethyldicocoammonium chloride to facilitate dispersion of the polymer in the media and improve suspension properties. The grafting reaction is performed as previously described using an appropriate redox system such as the ferrous salt with a source of peroxide. Since the metal ions are not removed from the product by washing as when a dry product is formed, a sequestrant for the metal ions may be added to the slurry at the conclusion of the reaction. The polymerization product has been found to remain readily dispersible or suspended in the slurry form over a period of time to facilitate storage and handling.

[0025] This method of grafting functions equally well for guar and hydroxypropyl guar to which vinyl phosphonic acid is grafted. The grafting conditions are as described above.

[0026] Graft copolymers of the present invention solubilize in aqueous liquids and substantially increase the viscosity of aqueous liquids. The viscosity of the copolymer solution may be further increased with the addition of a selected crosslinking activator or agent which effects an initiation of a crosslink interaction. Preferred crosslinking activators or agents in accordance with this invention comprise Bronsted-Lowry or Lewis bases which generally are substantially

free of polyvalent metal ions that is, metal ions having more than one valence state. Suitable compounds include, for example, calcium oxide, magnesium oxide and compounds selected from the group of mono, di and trialkanol amines such as triethanolamine, sodium hydroxide, potassium hydroxide, ammonia, various cocoamines such as Bis(2-hydroxyethyl) cocoamine, various pentamines such as tetraethylenepentamine, and various other water soluble amines, such as propyldiethanolamine, triethylamine, various water soluble borates such as the commercially available product Polybor, an admixture of boric acid and borate salts, from U.S. Borax and the like in the presence of a divalent cation, such as calcium or magnesium, which is present in at least a trace amount and which may be present in the aqueous liquid utilized to hydrate the copolymer or added as an additional component to the aqueous liquid. Such compounds generally are substantially free of polyvalent metal ions, that is, metal ions having more than one valence state. A particularly preferred crosslinking agent is magnesium oxide. As a result of the limited solubility of magnesium oxide in an aqueous solution, the rate of crosslink development is retarded or delayed such that a gelled fluid can be readily pumped into a wellbore for entry into a subterranean formation before significant crosslinking occurs in the fluid.

[0027] A surprising effect has been noted in the crosslinking of the polymer of the present invention in aqueous fluids containing salts. It has been found that when the polymer is hydrated in deionized water to which a small quantity of HCl has been added to facilitate hydration, the pH must be raised by the addition of the crosslinking activator to a level of about 6 before any significant crosslinking occurs. When the polymer is hydrated in an aqueous salt solution, such as for example CaCl_2 brine to which a small quantity of HCl is added, the pH at which crosslinking occurs is substantially lower. For example, in a 9 pound per gallon (1078 g/l) density CaCl_2 brine, crosslinking has been found to occur at a pH of about 4.5 and when a 13.5 pound per gallon (1618 g/l) density $\text{CaCl}_2/\text{CaBr}_2$ brine is employed, crosslinking was found to occur after the addition of sufficient crosslinking activator to raise the pH of the solution to a level of about 1.5.

[0028] While the specific mechanism by which the crosslinking occurs is unknown, it is believed that the crosslink is formed through the phosphorus moiety in the graft polymer which is activated by the presence of the Lewis or Bronsted-Lowry base.

[0029] While the following description will be directed to the use of magnesium oxide as a crosslinking activator or agent, it is to be understood that the described method would apply generally to any of the other suitable crosslinking activators of the present invention. A base gel is prepared by hydrating the previously described graft copolymer of hydroxyethyl cellulose, hydroxypropyl cellulose, guar or hydroxypropyl guar in an aqueous fluid at a pH in the range of from about 0.1 to about 3.5. The graft copolymer can be admixed with the aqueous fluid in an amount of from about 10 to about 300 pounds per 1000 gallons (1 to 36 g/l) of fluid. After the gelling agent has substantially hydrated, the base gel is admixed with a quantity of magnesium oxide.

[0030] The mixing can be effected in substantially any conventional mixing apparatus. The magnesium oxide generally is admixed with the base gel in an amount of from about 3 to about 30 pounds per 1000 gals (0.4 to 4 g/l) of base gel. In a preferred method of addition, the magnesium oxide is added as an aqueous slurry to facilitate handling of the material.

[0031] The fluid admixture containing the magnesium oxide then is introduced into a wellbore penetrating a subterranean formation. As the fluid is heated by passage down the wellbore the solubility of magnesium oxide in the fluid increases and the rate of crosslinking accelerates to form a crosslinked gelled fluid for treatment of the subterranean formation. When the fluid is to be used in, for example, a fracturing treatment, suitable particulates comprising any of the various materials known as propping agents may be added to the fluid. Similarly, if the fluid is to be used to perform a gravel pack operation, graded particulates may be added to the fluid before entry into the wellbore to form an appropriate gravel pack in a desired zone of a wellbore within a subterranean formation.

[0032] The fluid also can contain any other conventional additive such as gel stabilizers, breakers, clay stabilizers, bactericides, fluid loss additives and the like which do not adversely react with the fluid to prevent its use in a desired manner.

[0033] The following examples are provided to illustrate the utility of the composition of the present invention, however the invention is not to be considered as limited by these examples.

EXAMPLE I

[0034] To a 5 liter round bottom kettle, equipped with a stirrer, temperature controller and a N_2 sparge tube, the following reactants were added, about 2380 grams of PPG-400 (polypropylene glycol from Texaco Chemical Co.) and about 60 grams Claytone AF (alkyl quaternary ammonium montmorillonite from E. C. C. America, Inc.). The mix is stirred and nitrogen gas sparging is begun. Thereafter 1079 grams of hydroxyethyl cellulose (MS of about 2.2) is added and heating is initiated to slowly raise the temperature to about 40°C (generally 30 min to 1 hr.). After the temperature is reached, the temperature is maintained for about 1 hour to remove oxygen contamination.

[0035] While the above mixture is being heated about 319 grams of deionized water are admixed with about 10.5 grams of ferrous ammonium sulfate (reagent grade) in an erlenmeyer flask, while sparging, and dissolved. To this mixture is added about 121 grams of vinyl phosphonic acid from Hoechst Celanese Corporation and mixing and sparg-

ing is continued until the materials are dissolved. The solution then is added at the end of the sparging period to the 5 liter kettle.

[0036] The temperature is maintained while stirring and sparging and after about 1 hour 17.3 grams of 70% t-butyl-hydroperoxide is added to the kettle which then is allowed to react for about an hour. After the reaction is completed, a sequestrant, such as DEQUEST® 2010 from Monsanto Company, is added to the slurry to sequester the metal ions present and stirred. The reaction mixture then is permitted to cool. The reaction produced a 30% active polymer slurry.

[0037] A one (1) liter sample is prepared by mixing a sufficient quantity of the polymer with tap water to yield a 120 pound HEC/1000 gallon (14 g/f) solution. To facilitate hydration, sufficient acid, such as 15% hydrochloric acid, is admixed with the solution to correspond to about 10 gallons (38ℓ) per 1000 gallons (3785ℓ) of solution. The acid may be omitted or lesser quantities may be used if slower hydration is acceptable or desired. It is to be understood that other acid concentrations also could be utilized.

[0038] After being permitted to hydrate to form a gel, the gel is admixed with the equivalent of about 15 pounds per 1000 gallons (2 g/f) of magnesium oxide in the form of an aqueous slurry and a sample is evaluated with a FANN Model 35® viscometer. The pH of an aliquot sample also is monitored. The viscometer is equipped with a #1 spring, standard bob and sleeve. The results of the test are set forth below in Table I.

TABLE I

Time Minutes	Viscometer Dial Reading at 5.11 Sec ⁻¹	pH
1	18	1.98
5	18	5.28
6	18	5.70
7	19	5.98
8	20	6.17
9	22	6.30
10	23	6.41
11	25	6.50
12	27	6.57
13	29	6.63
15	33	6.74
20	51	6.92
25	70	7.06
30	93	7.17

[0039] The results clearly demonstrate the crosslink development upon addition of the base to the gelled fluid. The above test is repeated utilizing a 9 pound per gallon (1078 g/ℓ) density CaCl₂ brine instead of tap water. The results of the test are set forth in Table II, below.

TABLE II

Time Minutes	Viscometer Dial Reading at 5.11 Sec ⁻¹	pH
1	27	1.5
2	28	3.35
3	27	4.29
4	35	4.79
5	60	5.09
6	160	5.29
7	260	5.44

EXAMPLE II

[0040] To further demonstrate the crosslinkability of polymers produced in accordance with the present invention, the following test was performed. A polymer solution corresponding to 120 pounds per 1000 gallons (14 g/l) of 9 pound per gallon (1078 g/l) density CaCl_2 brine was prepared to which was added an equivalent of 28 pounds of triethanolamine per 1000 gallons (3 g/l) of solution. The fluid then was placed on a FANN Model 500 viscometer and evaluated at 176°F (80°C). The results of the test are set forth below in Table III.

TABLE III

Time After Addition Of Crosslink Activator, Minutes	Time At Temperature, Minutes	Fluid Temperature, F. (°C)	Apparent Viscosity, (cps) at 170 Sec ⁻¹
24	0	82 (28)	1087
28	4	177 (80.6)	919
43	19	180 (82)	924
58	34	177 (80.6)	947
73	49	176 (80)	910
88	64	176 (80)	902
103	79	176 (80)	995
118	94	176 (80)	991
133	109	176 (80)	952
148	124	176 (80)	973

[0041] The results of the elevated temperature test clearly demonstrate the stability of the crosslinked gel formed in accordance with the present invention.

EXAMPLE III

[0042] To further demonstrate the crosslinkability of the gelled fluids of the present invention with various bases, the following tests were performed.

[0043] A polymer solution corresponding to 120 pounds per 1000 gallons (14 g/l) of fluid was prepared. The aqueous fluid used was either 2% KCl solution prepared in water containing divalent cations or 9 pound per gallon (1078 g/l) density CaCl_2 brine. The polymer was hydrated by two different methods. Hydration method I comprised the addition of the gelling agent to a quantity of aqueous fluid followed by addition of further quantities of aqueous fluid during mixing. Hydration method II was the same as I with the addition of a quantity of 15% HCl equivalent to 10 gallons (38 l) per 1000 gallons (3785 l) of solution to accelerate the rate of hydration. Quantities of various bases then were admixed with the gel in order to determine the crosslinkability. The results are set forth in Table IV, below.

TABLE IV

Base	Quantity gallon (3.8 l) per 1000 gallon (37.85 l)	Aqueous Fluid	Original pH	Final pH	Gel Hydration Method	Comments
Triethanol Amine	2 (7.6 l)	9#/gal CaCl_2	1.98	4.5	II	Crosslinked
Triethanol Amine	2 (7.6 l)	9#/gal CaCl_2	3.05	6.04	I	Crosslinked
Triethanol Amine	2 (7.6 l)	2% KCl	2.62	7.03	II	Crosslinked
Tetraethylene- pentamine	2 (7.6 l)	9#/gal CaCl_2	3.0	8.5	I	Crosslinked
Bis (2-hydroxyethyl) cocamine	6 (22.7 l)	9#/gal CaCl_2	3.0	5.52	I	Crosslinked
Polybor	30#/1000 gallon	9#/gal CaCl_2	3.0	5.97	I	Crosslinked
None	0	9#/gal CaCl_2	3.0	3.0	I	No Crosslink
None	0	9#/gal CaCl_2	1.98	1.98	II	No Crosslink

EXAMPLE IV

[0044] To evaluate the effect of the crosslinked gel of the present invention upon a subterranean formation, the following regained permeability test was performed utilizing a Berea sandstone core sample. The test is performed at a fluid temperature of 150°F (66°C). The core sample is placed in the bottom of a vertical Hassler sleeve and a 1.9 cm hollow stainless steel spacer is positioned on top of the core. Annulus pressure of about 200 psi (about 1378952 Pa) above the highest head pressure (about 500 psi (about 3447380 Pa)) then is applied to the core. Initially, 250 ml of 15% HCl is flowed downwardly through the core at 50 psi (about 344738 Pa) to remove any carbonates from the core. Heated API brine then is flowed upwardly through the core at 20 psi (about 137895 Pa) until a stable flow rate is achieved to determine the initial core permeability. This step generally requires 800-1000 ml of fluid. A 200 ml sample of the crosslinked gel of the present invention then is prepared. Immediately after mixing the crosslinking activator with the gel, the gel is poured into a heated reservoir and fluid is flowed across the top of the core until the brine contained

in the spacer is fully displaced. The Hassler sleeve then is left fully shut-in for about 1 hour to ensure the gel is fully crosslinked. Thereafter, the valves on the Hassler sleeve are opened and 500 psi (about 3447380 Pa) pressure is applied to the reservoir containing the crosslinked fluid. This pressure is held for 2 hours and effluent from the bottom valve of the Hassler sleeve is collected to obtain an indication of fluid loss. The bottom valve then is closed and the 500 psi (about 3447380 Pa) pressure on the reservoir is released. A 15% HCl solution then is flowed across the core to clean out the mandrels and hollow spacer. After approximately 50 ml has passed across the top of the core, the Hassler sleeve is shut in for about 2 hours to permit the acid to break the gel. The valves are opened and a further 50 ml of 15% HCl is flowed across the core to flush any remaining residual gel from the mandrels and hollow spacer. Heated API brine then is flowed at 20 psi (about 137895 Pa) upwardly through the core sample until stable flow is achieved and the regained permeability is calculated.

[0045] The test fluid corresponded to the fluid of Example I prepared with 9 pound per gallon density (1078 g/l) CaCl_2 brine. The crosslinking activator was MgO admixed at a concentration of 15 pounds per 1000 gallons (2g/l) of fluid. The initial average API brine flow rate through the core was 11.55 ml per min.

[0046] The final average API brine flow rate was 11.42 ml per min. The regained permeability was calculated to be 99%. The results clearly demonstrate the substantially non-damaging behavior of the crosslinked gel of the present invention when placed in contact with a formation.

EXAMPLE V

[0047] To illustrate the crosslinkability of the grafted hydroxypropyl guar (HPG) copolymer of the present invention in comparison to ungrafted hydroxypropyl guar, the following tests were performed. Samples were prepared corresponding to a 120 pound per 1000 gallons (14 g/l) fluid with 9 pound per gallon (1078 g/l) CaCl_2 brine. The first sample contained ungrafted HPG, the second sample contained grafted HPG hydrated in the presence of the equivalent of 10 gallon (38 l) per 1000 gallons (3785 l) fluid of 15% HCl, the third sample contained grafted HPG hydrated without HCl addition. The grafting was effected as set forth in Example I and yielded a 30% active slurry of the copolymer. Samples No. 1 and 3 were admixed with MgO in an amount equivalent to 10 pounds per 1000 gallons (1 g/l) base gel in the form of an aqueous slurry. Sample No. 2 containing HCl was admixed with the equivalent of 15 pounds per 1000 gallons (2g/l) base gel. The results of the test are set forth in Table V, below.

TABLE V

Time, minutes after crosslink activator addition	Sample No. 1	Sample No. 2	Sample No. 3
4	no crosslinking	no crosslinking	no crosslinking
6	no crosslinking	1" lipping crosslink	2" lipping crosslink
8	no crosslinking	crosslinked non-pourable	crosslinked non-pourable
10	no crosslinking	crosslinked non-pourable	crosslinked non-pourable
15	no crosslinking	crosslinked non-pourable	crosslinked non-pourable
360	no crosslinking	crosslinked non-pourable	crosslinked non-pourable

EXAMPLE VI

[0048] To demonstrate the stability of the crosslinked gels of the present invention at elevated temperatures, the following test was performed. Samples were prepared with the aqueous fluids identified in Table VI, below. The gelling agent was that of Example I and was present in an amount equivalent to 120 pounds per 1000 gallons (14 g/l) of fluid. The equivalent of 10 gallons (38ℓ) per 1000 gallons (3785 ℓ) of fluid of 15% HCl was added to the gel to accelerate hydration. The quantity of MgO crosslinking activator added to each sample is set forth in Table VI. The samples then were placed in 12 ounce water (591 cm³) sample bottles identified as Citrate of Magnesia bottles. Each sample included 30 pounds per 1000 gallons (4 g/l) of fluid of sodium thiosulfate, a conventionally used gel stabilizer, to assist in stabilizing the fluid samples. After about 1 hour, a marble was placed on top of each of the crosslinked gel samples and the samples were placed in a 260°F. (127°C) oven. The stability of the gel was determined by settling of the marble in the gel sample upon heating and by visible changes in the gel. The results are set forth in Table VII, below.

TABLE VI

Sample No.	Aqueous Fluid	Crosslink Activator Quantity, Pounds Per 1000 gallon
1	10.5#/gal NaBr	20 (2.4g/l)
2	10.5#/gal NaBr	30 (4g/l)
3	8.6#/gal NaCl	20 (2.4g/l)
4	8.6#/gal NaCl	30 (4g/l)
5	10.0#/gal NaCl	20 (2.4g/l)
6	10.0#/gal NaCl	30 (4g/l)
7	3% KCl	20 (2.4g/l)
8	3% KCl	30 (4g/l)
9	Synthetic seawater	20 (2.4g/l)
10	Synthetic seawater	30 (4g/l)
11	Synthetic seawater	20 (2.4g/l)
12	Synthetic seawater + 3% KCl	30 (4g/l)

TABLE VII

Time, hours at 260°F (127°C)	S A M P L E N O ' S											
	1	2	3	4	5	6	7	8	9	10	11	12
24	Y	Y	Y	Y	F.W.	Y	Y	Y	Y	Y	Y	Y
68	Y	Y	Y	Y	W.R.	W.R.	Y	Y	Y	Y	Y	Y
96	L	Y	Y	Y	S	S	Y	Y	Y	Y	Y	Y

Y: = marble supported on gel surface

F.W.: = free water present in sample
bottle, marble on gel surface

W.R.: = water ring around gel in
sample bottle, marble on gel
surface

S: = about 30 - 50% free water present and
marble on gel surface

L: = sample bottle cap seal
leaked, no data collected

[0049] The foregoing test illustrates the ability of the gels of the present invention to carry a particulate such as proppant or a gravel pack material when exposed to elevated temperatures such as might occur in a subterranean formation.

Claims

1. A method of treating a subterranean formation, which method comprises contacting the formation with a treating fluid composition which comprises a mixture of (i) an aqueous liquid containing at least a trace amount of at least one divalent cation, (ii) a polymer derivative made by reacting one or more of the polymers guar, hydroxypropylguar, hydroxyethyl cellulose and hydroxypropyl cellulose, with a vinyl phosphonic acid in the presence of a redox system, and (iii) a crosslinking activator comprising a Lewis base or a Bronsted-Lowry base which is substantially free of polyvalent metal ions.
2. A method according to claim 1, wherein the polymer is reacted with a vinyl phosphonic acid in a reaction medium comprising at least one of tetramethyl ammonium chloride, polyethylene glycol and polypropylene glycol, and a redox initiator.
3. A method according to claim 2, wherein the reaction medium also includes a dispersant.
4. A method according to claim 3, wherein said dispersant comprises an alkyl quaternary ammonium montmorillonite, or dimethyldicocoammonium chloride, or any mixture of two or more thereof.
5. A method according to any of claims 1 to 4, wherein the reaction by which the polymer derivative is made is carried out at a temperature of from 20° to 60°C.
6. A method according to any of claims 1 to 5, wherein the crosslinking activator comprises magnesium oxide.

7. A method according to any of claims 1 to 5, wherein the crosslinking activator comprises at least one compound selected from magnesium oxide, a mono, di or trialkanol amine, calcium oxide, sodium hydroxide, potassium hydroxide, ammonia, a cocoamine, a pentamine, an alkyldiethanol amine, a mixture of boric acid and a borate salt, and diethylamine.
8. A method according to any of claims 1 to 7, wherein the redox system comprises a peroxide and a metal ion reductant.
9. A method according to claim 8, wherein the metal ion reductant comprises a source of ferrous ions.

Patentansprüche

1. Vorgehensweise des Treatings einer unterirdischen Formation, bestehend aus folgendem: dem Inberührungbringen der Formation mit einer Treatingflüssigkeit, bestehend aus einer Mischung aus (i) einer wässrigen Flüssigkeit, die wenigstens eine Spurmenge wenigstens eines zweiwertigen Kations aufweist, (ii) einem Polymerderivat, das durch die Reaktion von einem oder mehreren Polymerguar, Hydroxypropylguar, Hydroxyethylzellstoff und Hydroxypropylzellulose mit einer Vinylphosphonsäure bei Anwesenheit eines Redoxsystems reagiert wird sowie (iii) einem Vernetzungsaktivator, bestehend aus einer Lewis-Basis oder einer Bronsted-Lewis-Basis, der größtenteils frei von polyvalenten Metallionen ist.
2. Vorgehensweise nach Anspruch 1, bei der das Polymer mit einer Vinylphosphonsäure in einem Reaktionsmedium reagiert wird, das wenigstens aus Tetramethylammoniumchlorid, Polyethylenglykol und Polypropylenglykol sowie einem Redoxinitiator besteht.
3. Vorgehensweise nach Anspruch 2, bei der das Reaktionsmedium ebenfalls ein Dispersionsmittel enthält.
4. Vorgehensweise nach Anspruch 3, bei der das Dispersionsmittel aus einem Alkylquartärammonium-Montmorillonit oder Dimethyldicoco-Ammoniumchlorid oder einer Mischung von zwei oder mehreren dieser besteht.
5. Vorgehensweise nach einem der Ansprüche 1 bis 4, bei der die Reaktion, durch die das Polymerderivat erzeugt wird, bei einer Temperatur im Bereich von 20° bis 60° C erfolgt.
6. Vorgehensweise nach einem der Ansprüche 1 bis 5, wobei der Vernetzungsaktivator aus Magnesiumoxid besteht.
7. Vorgehensweise nach einem der Ansprüche 1 bis 5, wobei der Vernetzungsaktivator wenigstens aus einem Gemisch besteht, das aus Magnesiumoxid, einem Mono-, Di- oder Trialkanolamin, Kalziumoxid, Natriumhydroxid, Potassiumhydroxid, Ammoniak, einem Cocoamin, einem Pentamin, einem Alkyldiethanolamin, einer Mischung aus Borsäure und einem Borsalz und Diethylamin ausgewählt wird.
8. Vorgehensweise nach einem der Ansprüche 1 bis 7, bei der das Redoxsystem ein Peroxid und ein Metallionreduktionsprodukt umfaßt.
9. Vorgehensweise nach Anspruch 8, bei der das Metallionreduktionsprodukt eine Quelle von Eisenionen umfaßt.

Revendications

1. Un procédé de traitement d'une formation souterraine, ledit procédé comportant la mise en contact de la formation avec une composition de fluide de traitement qui se compose d'un mélange de (i) un liquide aqueux contenant au moins une trace d'au moins un cation bivalent, (ii) un dérivé polymère élaboré en faisant réagir un ou davantage des polymères guar, hydroxypropyl-guar, hydroxyéthylcellulose et hydroxypropylcellulose, avec un acide phosphonique vinylique en présence d'un système redox, et (iii) un agent activateur de réticulation comportant une base de Lewis ou une base de Bronsted-Lowry substantiellement exempte d'ions métalliques polyvalents.
2. Un procédé selon la revendication 1, selon lequel le polymère est soumis à une réaction avec un acide phosphonique vinylique dans un milieu de réaction comportant au moins l'un de: chlorure d'ammonium tétraméthyle, polyéthylèneglycol et polypropylèneglycol et un initiateur redox.

EP 0 604 115 B1

3. Un procédé selon la revendication 2, selon lequel le milieu de réaction contient également un agent de dispersion.
4. Un procédé selon la revendication 3, selon lequel ledit agent de dispersion comporte: montmorillonite d'ammonium quaternaire alkyle ou chlorure de diméthylidicocoammonium, ou un quelconque mélange de deux de ceux-ci ou
5 davantage
5. Un procédé selon l'une quelconque des revendications 1 à 4, selon lequel la réaction par laquelle le dérivé polymère est élaboré est réalisée à une température de 20°C à 60°C.
- 10 6. Un procédé selon l'une quelconque des revendications 1 à 5, selon lequel l'activateur de réticulation comporte un oxyde de magnésium.
7. Un procédé selon l'une quelconque des revendications 1 à 5, selon lequel l'activateur de réticulation comporte au moins un composé sélectionné parmi: oxyde de magnésium, mono-, di- ou trialkanolamine, oxyde de calcium,
15 hydroxyde de sodium, hydroxyde de potassium, ammoniac, une cocoamine, une pentamine, une alkyldiéthanolamine, un mélange d'acide borique et d'un sel de borate, et diéthylamine.
8. Un procédé selon l'une quelconque des revendications 1 à 7, selon lequel le système rédox comporte un peroxyde et un réducteur à ions métalliques
- 20 9. Un procédé selon la revendication 8 selon lequel le réducteur à ions métalliques comporte une source d'ions ferreux

